

ChemistrYES! B. Bonelli, M.U. Tomalino

Errata Corrige to the First Edition

Page 11: MnO₄²⁻ manganate ion

Page 14, Paragraph 2.4.3: H₂S_(aq) is hydrosulphuric acid, HCN_(aq) is hydrocyanic acid, HF_(aq) is hydrofluoric acid,

Page 21, eq. (10):

$$\text{electron mass} = \text{charge} \times \frac{\text{mass}}{\text{charge}} = (-1.602 \times 10^{-19}) (-5.686 \times 10^{-12} \text{ kg/C}) = 9.109 \times 10^{-31} \text{ kg} = 9.109 \times 10^{-28} \text{ g} \quad (10)$$

Page 73, eq (24):

|χ_A – χ_B| according to eq. (24):

$$|\chi_A - \chi_B| = c [BE_{AB} - (BE_{AA} * BE_{BB})^{\frac{1}{2}}] \quad (24)$$

Page 162, Table 10.3:

Compound	Intermolecular force	Superficial tension (J/m ²) at 20°C
Diethyl ether CH ₃ CH ₂ -O-CH ₂ CH ₃	Dipole-dipole; London	1.7*10 ⁻²
Ethanol CH ₃ CH ₂ -O-H	H bond	2.3*10 ⁻²
Water	H bond	7.3*10 ⁻²

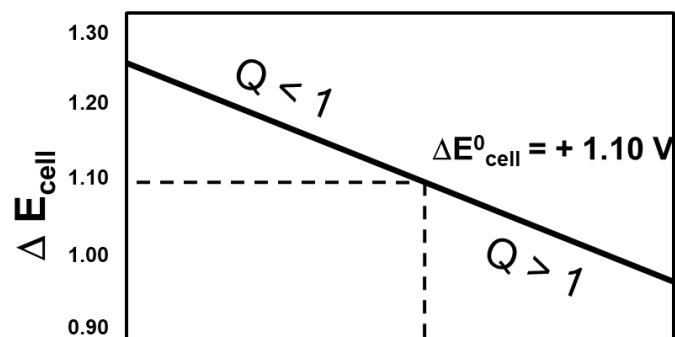
Page 225, Table 13.1

Disturbance	Reaction direction	Effect on Q	Effect on K(T)
[R] increases	right →	Q < K	None
[R] decreases	left ←	Q > K	None
[P] increases	left ←	Q > K	None
[P] decreases	right →	Q < K	None
For gas phase reactions with Δn = 0			
P increases (V decreases)	no shift	None	None
P decreases (V increases)	no shift	None	None
For gas phase reactions with Δn < 0			
P increases (V decreases)	right →	Q < K	None
P decreases (V increases)	left ←	Q > K	None
For gas phase reactions with Δn > 0			
P increase (V decreases)	left ←	Q > K	None
P decreases (V increases)	right →	Q < K	None

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Kinetic law	Zero order	First order	Second order
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Page 302, Figure 16.5



$$\log Q = [\text{Zn}^{2+}] / [\text{Cu}^{2+}]$$