Let's practice for the Chemistry exam: *errata corrige*

Page 37 (3.6 Practice yourself)

1) Calculate~~d~~ the moles…

9) Calculate the percentage purity of a raw mineral … with 120.0 g of ferrous sulphate and $1.00*10^2$ moles…

Page 40 (Exercise n° 4.1)

A given amount of gas occupies…

Page 45 (Exercise n° 4.8)

Sulphurous anhydride (SO_2) has M = 64.06 g mol^{-1}; chlorine (Cl_2) has mass 70.90 g mol^{-1}…

Page 46 (Exercise n° 4.9)

…Cl_2 has a molar mass M of 70.90 g mol^{-1}…

Page 51 (Exercise n° 5.1)

1) $S_{(s)} + 3/2\ O_{2(g)} \rightarrow SO_{3(g)}$

Lesson 6: the abbreviation *sz* must be replaced by *sn*

Page 70 (6.3 Practice yourself)

1) …1537. mL of water at 25.0…
2) …15.0 wt% ethylene glycol (k_b of water is 0.512 °C kg mol^{-1})

4) … HCl solution, $2.00*10^2$ mL

Page 71 (6.3 Practice yourself)

7) (P^0 = 18.7 mmHg at 21.0 °C)

11) Determine the molecular mass and the molecular formula…

Page 87 (Exercise 12)

…and 0.500 moles of CO_2.

Page 93 (Exercise 8.5)

…that volumes are additive~~s~~

Page 107 ΔE_{cell}~~a~~

Page 110 (Exercise n° 9.4)

$E^0_{Ni2+/Ni} = -0.23\sout{3}5\ V$

Page 123 (Exercize 14)

During electrolysis of an aqueous solution of cupric sulphate…

Page 129 (**Exercise 10.11**): $CH_3COOCH(CH_3)_2$ (isopropyl ~~formate~~ acetate)